

## TITLE OF THE INVENTION

SOLID POLYMER ELECTROLYTE FUEL CELL AND  
METHOD FOR PRODUCING ELECTRODE THEREOF

## CROSS-REFERENCE TO RELATED APPLICATIONS

[0001] This application is based on and claims priority under 35 U.S.C. §119 with respect to Japanese Patent Application No. 2000-355722 filed on November 22, 2000, the entire content of which is incorporated herein by reference.

## BACKGROUND OF THE INVENTION

## Field of the Invention

[0002] This invention relates to a solid polymer electrolyte fuel cell and a method for producing an electrode of the polymer electrolyte.

## Discussion of the Background

[0003] Conventional solid polymer electrolyte fuel cells include a proton-conductive solid polymer electrolyte membrane. Solid polymer electrolyte fuel cells generate electromotive force by electrochemical reactions between fuel gas (for example, H<sub>2</sub> gas) and oxidizer gas.

[0004] The solid polymer electrolyte fuel cells produce the following electrochemical reaction between H<sub>2</sub> gas as the fuel gas and the oxidizer gas as the O<sub>2</sub> gas at an anode side.



[0005] After the electrochemical reaction, the resulting proton (H<sup>+</sup>) passes through the solid polymer electrolyte membrane. Then the solid polymer electrolyte fuel cell produces the next reaction at a cathode.



[0006] Accordingly, the electrolyte fuel cell produces an electromotive force between the anode and cathodes of up to 1.23 V.

[0007] The schematic drawing of the conventional solid polymer electrolyte fuel cell is shown in Fig. 3. To smoothly and efficiently produce the above electrochemical reactions of the fuel cell shown in the reaction formulas 1 and 2, gas diffusion electrode 1b in Fig. 3 plays

a very important role.

**[0008]** For the solid polymer electrolyte fuel cell to generate electric power, each of the fuel gas and the oxidizer gas needs to be supplied to surfaces of catalytic layers 1d disposed on the electrodes. At the cathode, water is generated on the surface of the catalytic layer 1d, as shown in the reaction formula (2). This water covers the surface of the catalytic layer 1d to inhibit the oxidizer gas from being supplied to the catalytic surface 1d.

**[0009]** In the anode 1bb protons ( $H^+$ ) generated by the reaction as shown in the reaction formula (1) hydrate or take the water to the cathode 1ba through the solid polymer electrolyte membrane 1a. But the water in the anode 1bb is absorbed and stops short. Therefore the solid polymer electrolyte membrane 1a becomes dried out. To prevent the catalytic layer 1d from being dried, the fuel gas supplied to the catalytic layer 1d at the anode 1bb is generally humidified. But the excessive humidification by the fuel gas inhibits the fuel gas from being supplied to the catalytic surface 1d at the anode. To avoid the flooding due to water generated by the reactions and the water added by the humidified fuel gas, the electrode has been mixed with a water-repellent 1c, i.e., Polytetrafluoroethylene (PTFE) and so on. But the excessive addition of the water-repellent 1c to the electrodes 1b increases the electric resistance both in the bulk and on the surface of the electrodes 1b and decreases the gas permeability of the electrodes 1b. The excessive hydrophobicity of the electrodes 1b inhibits the humidification of the solid polymer electrolyte membrane 1a. Furthermore, the water absorbed in the solid polymer electrolyte membrane 1a is taken away by the fuel gas and the oxidizer gas provided to the catalyst layer 1d. Therefore the solid polymer electrolyte membrane 1a is dried out.

**[0010]** In order to improve the efficiency of the catalyst activity by three-dimensionally using the catalyst layer 1d, the electrolyte of the electrolyte membrane 1a has been dissolved with a catalyst in a solvent and impregnated with a gas diffusion layer wherein the gas can be diffused. In this process, the electrodes acquire proton-conductivity and even hydrophilicity. The electrode 1b as a gas diffusion type electrode, i.e., the electrode formed with the gas diffusion layer, needs to easily and equally distribute the fuel gas and oxidizer gas to the surface of the catalyst layer 1d. For this to occur, and to increase the porosity content, the gas permeability and gas diffusion coefficient needs to be increased.

**[0011]** However, the excessive provisions of fuel gas and oxidizer gas causes the electrolyte membrane 1a to be dried out, so that the proton-conductivity of the solid polymer electrolyte

membrane 1a is reduced.

**[0012]** In addition, the conventional separators 1e of the fuel cell as shown in Fig. 3 generally are formed with concave portions and convex portions relative to the electrodes, in sectional shape, in order to output electric current and supply the fuel gas and the oxidizer gas to the electrodes 1bb and 1ba, respectively. Since the fuel gas and the oxidizer gas pass through the concave portions of the separators to be provided to the electrodes, but not at the convex portions, the fuel gas and the oxidizer gas are not equally diffused or distributed over the whole surface of each separator 1e. The electrodes 1b (the gas diffusion type electrodes) must diffuse the fuel gas or the oxidizer gas from the concave portions to the convex portions of the surface of the separators in order that the concentration of the fuel gas or the oxidizer gas is equally diffused in the surface of the catalyst layer 1d. For the above function, the electrodes 1b (the gas diffusion type electrodes) are made of a material having a large gas diffusion coefficient (porous material). But, the above material (porous electrode) tends to remove water from the electrolyte membrane 1a, to dry the solid polymer electrolyte membrane 1a.

**[0013]** As described above, it is necessary for the gas diffusion layer of the fuel cell to be made with an appropriate balance between hydrophilicity and hydrophobicity, and an appropriate gas permeability.

**[0014]** To form the electrode having hydrophilicity, hydrophobicity and gas permeability, the conventional electrode is by from the following process. First, carbon black CB and PTFE formed in a paste with a dispersion medium are shaped in a sheet form. Then the sheet is baked to sinter the PTFE. Alternatively, the CB and PTFE are sometimes impregnated with carbon cloth or carbon paper, then the impregnated sheet is baked.

**[0015]** Though the structure or content of the above electrode is decided by a complex agglutination/dispersion mechanism which is changed by types, contents and mixing methods of the carbon blacks CB, PTFEs and dispersion mediums, the internal structure of the electrode cannot be regulated depending on the designer's intent.

#### SUMMARY OF THE INVENTION

**[0016]** According to an aspect of the present invention, a solid polymer electrolyte fuel cell includes a polymer electrolyte membrane having proton-conductivity, an anode disposed on one surface of the polymer electrolyte membrane, and a cathode disposed on the other surface

of the polymer electrolyte membrane. The cathode has a first gas diffusion layer joined to a second gas diffusion layer in the thickness direction of the cathode. The second gas diffusion layer has different characteristics as compared to the first gas diffusion layer.

[0017] According to another aspect of the present invention, the cathode has a hydrophilic intermediate layer disposed between the first gas diffusion layer and the second gas diffusion layer of the cathode for holding water. The first gas diffusion layer differs from the second gas diffusion layer in any one of gas permeability in the thickness direction of the cathode, electric resistance and hydrophobicity. The first gas diffusion layer is disposed at a catalyst layer side of the cathode and the second gas diffusion layer is disposed at a separator side of the cathode. The gas permeability of the second gas diffusion layer is larger than the gas permeability of the first gas diffusion layer.

[0018] Another aspect of the invention involves a method for producing an electrode of the solid polymer electrolyte fuel cell and includes a first step of forming a plurality of second gas diffusion layers, a second step of forming a first gas diffusion layer by the pressing one of the second gas diffusion layers to increase a density of the second gas diffusion layer, and a third step of thermally pressing to join the first gas diffusion layer to the second gas diffusion layer.

#### BRIEF DESCRIPTION OF THE DRAWINGS

[0019] The foregoing and additional features and characteristics of the present invention will become more apparent from the following detailed description considered with reference to accompanying drawings in which like reference numerals designate like elements and wherein:

[0020] Fig. 1 shows a cross sectional view of a cathode according to the present invention;

[0021] Fig. 2 shows graphs for evaluation of a contacting diffusing layer and a reference diffusing layer; and

[0022] Fig. 3 shows a cross sectional view of a conventional solid polymer electrolyte fuel cell.

#### DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENT

[0023] Hereinafter, an embodiment of the present invention will be described with reference to the attached drawings. Requirements of a cathode of a solid polymer electrolyte fuel cell

are resistance to flooding by water in a catalyst and a necessity or a small necessity of humidification of fuel gas and/or oxidizer gas from outside of the fuel cell.

**[0024]** Fig. 1 shows a cathode 100 according to the present invention. The cathode (joined diffusion layer) is constructed by disposing a first gas diffusion layer 10 over a second diffusing layer 20, as a multiple layer structure. The cathode 100 includes a first gas diffusion layer 10 disposed at a catalyst layer 40 side and a second gas diffusion layer 20 disposed at a separator (not shown) side. The first gas diffusion layer 10 has small gas permeability compared to the second diffusing layer 20. The first gas diffusion layer 10 is made of a high density or airtight material layer. While the second gas diffusion layer 20 is made of a porous material layer which has larger porosity compared to that of the first gas diffusion layer 10.

**[0025]** The first gas diffusion layer 10 is produced by compressing the second gas diffusion layer 20 in the thickness direction thereof. Thus the first gas diffusion layer 10 will have a higher density compared to the second gas diffusion layer 20.

**[0026]** The cathode 100 is further provided with an intermediate layer 30 disposed between the first gas diffusion layer 10 and the second gas diffusion layer 20 for keeping or holding water. The intermediate layer 30 is made of a hydrophilic material, for example, carbon black CB and so on. The intermediate layer 30 absorbs and holds the water generated by an electrochemical reaction. Therefore the intermediate layer 30 prevents the water flooding the catalyst layer at the cathode side.

**[0027]** When the dried oxidizer gas reaches to the intermediate layer 30, a part of the dried oxidizer gas is diverted to take the water to outside of the cathode 100. While the remaining oxidizer gas passes through the intermediate layer 30 to be humidified, and is provided to the electrolyte membrane. By providing the above two separate passages for the oxidizer gas, the excessive humidification of the oxidizer gas is prevented.

**[0028]** In the cathode 100, the first gas diffusion layer 10 may differ from the second gas diffusion layer 20 in one or more of gas permeability in the thickness direction of the electrode, electrical resistance and hydrophobicity..

**[0029]** In accordance with the present invention, since the plural gas diffusion layers are connected or each other or joined therewith under pressure to form an electrode as described above, the structure of the electrode in the thickness direction (thickness, the characteristics of material and so on) can be readily controlled. By inserting other sheet made of material

having different characteristics in water-attraction, water-repellence and so on, the performance of the intermediate layers of the electrode can be separately changed.

**[0030]** As the present invention, joining plural gas diffusion layers which have the same or different characteristics provides following effects.

**[0031]** 1. Freedom in designing the electrode is dramatically improved.

**[0032]** 2. Since a combination of the gas diffusion layers can be flexibly changed, the specifications of the electrolyte membranes, separators, operating condition, and so on can be flexibly changed, so that a small stock of gas diffusion layers can be used for many types of electrodes. For example, a change of the thickness of the electrode can be provided by changing the number of gas diffusion layers to be joined under pressure.

**[0033]** 3. Since the plural gas diffusion layers which have same or different characteristics or sizes can be partially combined or joined, the cross sectional shape of the electrode or partial characteristics of the electrode can be changed.

**[0034]** 4. In addition, the electric resistance of the first gas diffusion layer may differ from that of the second gas diffusion layer. In this case, the electric correcting performance of the electrode will be improved.

## A. PRODUCTION OF ELECTRODE

### EXAMPLE

**[0035]** To produce the electrode shown in Fig. 1, an electrode (example) has been manufactured by the following steps. The gas diffusion layers to be joined are manufactured based on a method for producing a carbon paper disclosed in Japanese Patent Application Publication published on May 5, 2000 as Toku-Kai-Hei 2000 - 136493. Namely, the gas diffusion layers are formed in a wet papermaking process.

**[0037]** 1. Carbon fiber (diameter:  $\phi = 12.5 \mu\text{m}$ , length:  $L = 3 \text{ mm}$ ) and wood pulp are prepared in proportion 6 part to 4 in weight. The carbon fiber and the wood pulp are dispersed in water with an appropriate reagent to form a paste. The wood pulp is added as binder (fiber for binding the carbon fiber) to bind the carbon fibers so that the gas diffusion layer can be thin.

**[0038]** 2. The carbon fiber and pulp paste are formed to be a carbon/pulp paper (weight:  $50 \text{ g/m}^2$ , thickness:  $T = 0.2 \text{ mm}$ ) by means of the wet papermaking process by hand.

**[0039]** 3. Carbon black CB, ethylene glycol and isopropyl alcohol IPA are mixed to form a

paste. The paste is homogeneously impregnated into the carbon fiber/pulp paper using a screen printer. After the above impregnation, the ethylene glycol and isopropyl alcohol are removed from the impregnated carbon fiber/pulp paper by being heated at 85°C in the atmosphere until the carbon fiber/pulp paper is dried out.

[0040] 4. The PTFE is impregnated into the carbon fiber/pulp paper by dipping the carbon fiber/pulp paper into PTFE 15 wt% (weight percent) dispersion solution, for example, "Polyflon D1" manufactured by Daikin Industries, Ltd. After the above dipping, the impregnated carbon fiber/pulp paper is baked at 390°C in the atmosphere for 1 hour, until the PTFE is melted or dissolved and fitted to the carbon fiber paper. By the above baking, the wood pulp is burned to be ash, which is left in the vacancies in the baked carbon fiber/pulp paper. The resulting vacancies form gas communicating pores or vents in the electrode. Hereinafter, the gas diffusion layers manufactured by the above processes will be called the second gas diffusion layers 20.

[0041] 5. For controlling the gas permeability, the gas diffusion layer 20 is pressed in a press machine (not shown) to increase their densities. Hereinafter, the gas diffusion layer produced by pressing the second gas diffusion layer 20 at room temperature will be called the first gas diffusion layer 10.

[0042] 6. To form the hydrophilic layer in the intermediate portion of the electrodes in the thickness direction, the carbon black CB is impregnated on only one surface of the second gas diffusion layer 20. (Then the impregnated carbon black CB will be the intermediate layer 30 formed between the first gas diffusion layer 10 and the second gas diffusion layer 20 after the next process.)

[0043] 7. The above impregnated second gas diffusion layer 20 is joined to the first gas diffusion layer 10 by being pressed or hot-pressed by 80 kg/m<sup>2</sup> at 160°C for 1.5 minutes. Accordingly, the above multiplied gas diffusion layer is a joined diffusion layer 100 of the present invention.

#### REFERENCE SAMPLE (CONVENTIONAL ART)

[0044] 1. As in the example of the present invention, carbon fiber (diameter:  $\phi = 12.5 \mu\text{m}$ , length:  $L = 3 \text{ mm}$ ) and wood pulp are prepared in proportion 6 part to 4 in weight. The carbon fiber and pulp are dispersed in water with an appropriate reagent and formed in a paste.

[0045] 2. The carbon fiber and pulp paste are formed to be a carbon/pulp paper (weight: 50

g/m<sup>2</sup>, thickness: T = 0.4 mm).

[0046] 3. Carbon black CB, ethylene glycol and isopropyl alcohol (IPA) are mixed to be formed in a paste. The paste is equally impregnated into both surfaces of the carbon fiber/pulp paper by a screen printer. After the above impregnation, the ethylene glycol and isopropyl alcohol are removed from the impregnated carbon fiber/pulp paper by heating in 85°C atmosphere, until the carbon fiber/pulp paper is dried.

[0047] 4. The PTFE is impregnated into the carbon fiber/pulp paper by dipping the carbon fiber/pulp paper into PTFE 15 wt% (weight percent) dispersion solution, for example, "Polyflon" D1 manufactured by Daikin Industries, Ltd. After the above impregnation, the impregnated carbon fiber/pulp paper is baked at 390°C in atmosphere for 1 hour, until the PTFE is melted or dissolved and fitted to the carbon fiber paper.

[0048] 5. The above impregnated carbon fiber/pulp paper is pressed at room temperature to the same thickness as the joined diffusion layer 100. Accordingly the above cold-pressed carbon fiber paper is the diffusion layer to be compared as a reference sample.

#### METHOD FOR PRODUCING MEMBRANE-ELECTRODE ASSEMBLY

[0049] To measure the fuel cell electrode of the present invention, a membrane-electrode assembly, which is constructed by joining the solid electrolyte membrane to the electrode, is manufactured by following processes.

[0050] 1. Platinum supported by carbon black, polymer electrolyte solution (for example, "Aciplex solution", a trademark, manufactured by Asahi Chemical Industry Co., Ltd.), water and isopropyl alcohol are mixed in proportion 1:15 : 2:2 in weight to form a paste.

[0051] 2. The above catalyst paste is coated on one surface of the joined diffusion layer 100 at the first gas diffusion layer 10 side to a thickness of 300 nm. The catalyst paste is also coated on one side of the reference diffusion layer to a thickness of 300 nm. Both catalyst pastes are dried at 80°C in a vacuum.

[0052] 3. The joined diffusion layer 100 is cut into a circle which has a diameter  $\phi = 35.7$  and an area  $S = 10 \text{ cm}^2$ . The reference diffusion layer is cut into three circles which have the same diameter and the same area.

[0053] 4. The joined diffusion layer 100 and the reference diffusion layers cut into circles are joined with polymer electrolyte membranes (for example, "Nafion 112", a trade mark, manufactured by E.I. du Pont de Nemours and Company) by hot-pressing. Accordingly, a



membrane-electrode assembly is formed.

**[0054]** As two combinations of the electrodes, the following two membrane-electrode assemblies are manufactured.

**[0055]** Anode/Cathode = Reference gas diffusion layer/Joined diffusion layer

**[0056]** Anode/Cathode = Reference diffusion layer/Reference diffusion layer

**[0057]** The effects of the present invention are verified in application to the cathode.

## B. CHARACTERISTIC EVALUATION

**[0058]** Table 1 shows each thickness, gas permeability and bulk electric resistance of the electrodes.

**[0059]** The joined diffusion layer scarcely differs from the reference diffusion layer in the characteristics (thickness, gas permeability, and bulk electric resistance).

Table 1 - Gas Diffusion Layers Characteristics

	Diffusion layer 1	Diffusion layer 2	Joined diffusion layer	Reference diffusion layer
Thickness [mm] (under pressure at 20 kgf/cm <sup>2</sup> )	0.15	0.1	0.26	0.26
Gas permeability coefficient [ $\mu\text{m (Pa}\cdot\text{s)}^{-1}$ ]	450	35	32	30
Electric resistance (under pressure at 20 kgf/cm <sup>2</sup> )	18	8	24	23

**[0060]** Next, the joined diffusion layer and the reference diffusion layer were disposed at the cathode sides in the actual solid polymer electrolyte fuel cells, respectively. Then the difference in characteristics between the fuel cell with the joined diffusion layer and the fuel cell with the reference diffusion layer was evaluated when the fuel cells were actually operated. In these case, purified hydrogen gas ( $\text{H}_2$ ) is employed as the fuel gas, while, air was used as the oxidizer gas. Each gas was humidified by passing the gas through thermally controlled water. Thus each gas included a saturated vapor corresponding to the water temperature and was inducted into the fuel cell. In this evaluation, three temperatures of the water in bubbling at the cathode side were determined, the potential-current characteristic (V-I characteristic) and cell resistance were measured.

**[0061]** Operating conditions of the fuel cell were determined as follows.

[0062] Gas Pressure: Purified hydrogen/Air = 2 atm/2 atm

[0063] Stoichiometrical Ratio: Purified hydrogen/Air = 1.2/4

[0064] Cell Temperature: 80°C

[0065] Water temperature: Purified hydrogen/Air = 90°C/(a) 50°C, (b) 80°C, (c) 30°C

Table. 2 Cell Resistance with Respect to Each Gas Diffusion Layer

Water Temperature	Electric Resistance of Joined Diffusion layer [ $\text{m}\Omega\text{cm}^2$ ]	Electric Resistance of Reference Diffusion Layer [ $\text{m}\Omega\text{cm}^2$ ]
(a) 50°C	142	141
(b) 80°C	139	137
(c) 30°C	144	233

[0066] As shown in Table 2, if either the joined diffusion layer or the reference diffusion layer is employed in the cathode, the difference of the characteristics, i.e., a flooding-resistance and a dry-up-resistance, was verified.

[0067] Fig. 2 shows the differences in potential-current characteristics (V-I characteristics) between the example and the reference sample relative to the three water temperatures for the humidification for the gases.

[0068] Fig. 2-(a) especially shows the difference in the potential-current characteristics of the solid polymer electrolyte fuel cell when the water temperature of the humidification was set at 50°C. Even if any one of the joined diffusion layer and the reference diffusion layer was employed in the fuel cell, the characteristics of the fuel cells scarcely differ from each other.

[0069] Fig. 2-(b) shows the difference in the potential-current characteristics when the water temperature of the humidification was set at 80°C in order that the air was more humid than at 50°C. The characteristic of the cathode disposed with the joined diffusion layer was almost the same as at 50°C. To the contrary, according to the cathode disposed with the reference diffusion layer, the potential rapidly dropped relative to the current density area more than  $1 \text{ A/cm}^2$ . The drop of the potential in the fuel cell disposed with the referential layer was due to the flooding by the generated water in the catalyst layer. The cause of the flooding was assumed to be that the excess water quantity limits the water elimination performance of the reference diffusion layer.

[0070] Fig. 2-(c) shows the potential-current characteristics when the water temperature was set at 30°C in order that the humidification of the air at the cathode side was reduced to a low



